## A New Reaction Involving 1,5-Diazabicyclo[4.3.0]non-5-ene as a Nucleophile and a Two Proton Donor

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1,5-Diazabicyclo[4.3.0]non-5-ene (DBN) reacts at both phosphorus atoms of the (phosphino)(*P*-chlorophosphonio)carbene **1a** affording the cationic unsaturated tricyclic adduct **5**.

DBN (1,5-diazabicyclo[4.3.0]non-5-ene) is frequently used as a 'non-nucleophilic base' both in organic and inorganic syntheses.<sup>1</sup> However, it has been shown that this strong non-ionic base can react with a variety of main group electrophiles, including those which are relatively weak, giving rise to the formation of a range of onio-substituted derivatives.<sup>2</sup> We now report a further remarkable reaction in which DBN acts not only as a nucleophile but also as a two proton donor.

We have recently described the synthesis and spectroscopic characterisation in solution of the (phosphino)(*P*-chlorophosphonio)carbene 1a;<sup>3</sup> however all attempts to isolate this compound failed. *Ab initio* calculations<sup>4</sup> and an X-ray crystal



Scheme 1 (Phosphino)(phosphonio)carbenes 1a X = Cl, 1b X = H and the expected adduct 2 of 1a with DBN; counter anion  $CF_3SO_3^-$ ;  $R = NPri_2$ 



Fig. 1 Molecular structure of 5. Selected bond lengths (Å) and angles (°): P(1)–C(1) 1.660(2), P(2)–C(1) 1.719(3), P(1)–N(3) 1.654(2), P(1)–N(4) 1.636(2), P(2)–N(5) 1.666(2), P(2)–N(6) 1.685(3), P(1)–N(1) 1.705(2), P(2)–C(2) 1.751(3), N(1)–C(3) 1.384(4), C(2)–C(3) 1.358(4), C(3)–N(2) 1.387(4) Å; P(1)–C(1)–P(2) 123.6(2), C(1)–P(2)–C(2) 104.7(1), P(2)–C(2)–C(3) 123.0(2), C(3)–N(1)–P(1) 117.5(2), N(1)–P(1)–C(1) 108.7(1)°.

study<sup>5</sup> of the *P*-hydrogeno analogue **1b** have shown that the  $\sigma^3$ -phosphorus atom was planar and positively charged (Scheme 1). Therefore, by analogy with the reaction of DBU or DBN with phosphenium salts,<sup>2</sup> we believed it would be possible to isolate adduct **2**.

To a dichloromethane solution of carbene **1a** was added an excess of DBN at -78 °C. The quantitative conversion to a new species was indicated by <sup>31</sup>P NMR [AB system at  $\delta$  +42.5 and +58.9 (<sup>2</sup>J<sub>PP</sub> = 50.1 Hz)]. However, the presence of a PC(H)=P sequence was indicated by a *J*-modulated <sup>13</sup>C NMR spectrum (15.1, dd, <sup>1</sup>J<sub>PC</sub> = 179.4 and 138.2 Hz) and was confirmed by a doublet of doublets in the <sup>1</sup>H NMR spectrum [0.97 (<sup>2</sup>J<sub>PH</sub> = 0.9 and 11.2 Hz)]. Furthermore doublets at  $\delta$  68.9 (<sup>1</sup>J<sub>PC</sub> = 151.1 Hz) and 154.4 (<sup>2</sup>J<sub>PC</sub> = 11.6 Hz) in the <sup>13</sup>C NMR spectrum were consistent with a PC=C fragment.<sup>†</sup>

Crystals suitable for a single-crystal X-ray diffraction study‡ were isolated in 50% yield after washing the crude reaction mixture with ether and recrystallisation from  $CH_2Cl_2$ -ether. This study revealed that instead of the expected simple carbene adduct **2**, a tricyclic species **5** had been formed (Fig. 1). There is a weak interaction of the triflate anion with the methylene proton [C(1)…O(1) = 3.253(4) Å, C(1)-H(C1)…O(1) = 153°], but not with either of the phosphorus centres. The six membered ring containing the P-C(H)-P fragment adopts a boat-like conformation. The bond distances P(1)-C(1) [1.660(2) Å] and P(2)-C(1) [1.719(3) Å] are comparable, indicating partial charge delocalisation. A localised carbon-carbon double bond



Scheme 2 Proposed mechanism for the synthesis of the tricyclic adduct 5;  $R = NPr_{i_2}^{i_2}$ ; counter anion CF<sub>3</sub>SO<sub>3</sub><sup>-</sup>



C(2)–C(3) [1.358(4) Å] is readily apparent in the fivemembered ring.

This reaction is believed to proceed as expected *via* initial nucleophilic attack by DBN at the  $\sigma^3$ -phosphorus centre of **1a** resulting in the formation of **2**. The basicity of the carbodiphosphorane carbon is sufficient to abstract a relatively acidic proton in the  $\alpha$  position to the iminium function giving **3**. Cyclisation then occurs by attack of the enamine on the phosphonio centre affording **4**; lastly excess DBN easily removes HCl yielding product **5** (Scheme 2). Overall the reaction can be considered as the 1,3-di-nucleophilic attack of DBN on the (phosphino)(phosphonio)carbene **1a**. This type of reactivity is not unprecedented, Chambers *et al.* have recently reported that the related base DBU reacts in a similar manner with 2-H heptafluorobut-2-ene.<sup>6</sup>

Since very few cyclic carbodiphosphoranes are known,<sup>7</sup> it was tempting to try to deprotonate **5** (Scheme 3). Initial attempts with classical bases proved unsuccessful, so a procedure analogous to that described by Bestmann<sup>8</sup> was used. Compound **5** was treated with a slight excess of sodium hexamethyldisilyl amide in benzene, the <sup>31</sup>P NMR spectrum clearly demonstrated the formation of a new species (AB system at  $\delta$  +30.2 and +29.4, <sup>2</sup>*J*<sub>PP</sub> = 114.0 Hz). The deprotonation of **5** was apparent from the <sup>13</sup>C proton coupled NMR spectrum where the signal of the PCP carbon appears as a doublet of doublets (14.9, <sup>1</sup>*J*<sub>PP</sub> = 138.0 and 144.0 Hz). Compound **6** is easily reprotonated suggesting the highly basic nature of this new carbodiphosphorane, despite complete charge delocalisation in the conjugate acid **5** being precluded by the non-planar nature of the ring system.

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## Footnotes

† Satisfactory elemental analyses for 5 have been obtained.

 $\ddagger$  Crystal data for 5: C<sub>33</sub>H<sub>67</sub>N<sub>6</sub>F<sub>3</sub>P<sub>2</sub>S<sub>1</sub>O<sub>3</sub>, M = 746.9, monoclinic, space group  $P2_1/c$  (No. 14), a = 10.933(2), b = 9.889(2), c = 37.426(4) Å,  $\beta = 10.933(2)$ , b = 10.933(2), b = 10.933(2 $91.19(1)^{\circ}$ , V = 4045(1) Å<sup>3</sup>, Z = 4, F(000) = 1616,  $\mu$ (Mo-K $\alpha$ ) = 0.20  $mm^{-1}$ ,  $D_c = 1.226 \text{ g cm}^{-3}$ , crystal size  $0.50 \times 0.40 \times 0.40 \text{ mm}$ . 8801 total reflections were collected on an Enraf-Nonius CAD4 diffractometer at 293 K (graphite-monochromated Mo-K $\alpha$  radiation  $\lambda = 0.71073$  Å,  $\omega$ -2 $\theta$  scan mode,  $2\theta \le 54^\circ$ ). Of these, 5141 reflections with  $F_0^2 > 3\sigma(F_0^2)$  were used for the solution (direct methods using SHELXS-86 program) and refined by full-matrix least squares analysis (using SHELX-76 program). All non-H atoms anisotropic. H atoms found on a difference Fourier map, introduced in calculations in constrained geometry [d(C-H) = 0.97 Å], with isotropic thermal parameters first refined, then kept fixed. Scattering factors (f', f'')taken from 'International Tables for X-ray Crystallography.' Last fullmatrix refinement cycle:  $R(f_0) = 0.040$ ,  $R_w = 0.044$  with unit weights, 5141 observations, 433 variable parameters, S = 1.388, max. parameter shift/esd = 0.001, max. and min. residual peaks of 0.29 and -0.26 e Å<sup>-3</sup>. Atomic coordinates, bond lengths and angles, and thermal parameters have been deposited at the Cambridge Crystallographic Data Centre. See Information for Authors, Issue No. 1.

## References

- 1 H. Oediger, F. Möller and K. Eiter, Synthesis, 1972, 591.
- 2 R. Reed, R. Réau, F. Dahan and G. Bertrand, Angew. Chem., Int. Ed. Engl., 1993, 32, 399; G. Bouhadir, R. W. Reed, R. Réau and G. Bertrand, Heteroatom Chem., 1995, 6, 371.
- 3 P. Dyer, A. Baceiredo and G. Bertrand, Inorg. Chem., in the press.
- 4 O. Treutler, R. Ahlrichs and M. Soleilhavoup, J. Am. Chem. Soc., 1993, 115, 8788.
- 5 M. Soleilhavoup, A. Baceiredo, O. Treutler, R. Ahlrichs, M. Nieger and G. Bertrand, J. Am. Chem. Soc., 1992, **114**, 10959.
- 6 R. D. Chambers, A. J. Roche, A. S. Batsanov and J. A. K. Howard, J. Chem. Soc., Chem. Commun., 1994, 2055.
- 7 H. Schmidbaur, T. Costa, B. Milewski-Mahrla and B. Schubert, Angew. Chem., Int. Ed. Engl., 1980, 19, 555.
- 8 H. J. Bestmann, D. Hadawi, H. Behl, M. Bremer and F. Hampel, Angew. Chem., Int. Ed. Engl., 1993, 32, 1205.